Taming C₆₀ fullerene: tuning intramolecular photoinduced electron transfer process with subphthalocyanulines†

Marc Rudolf,a Olga Trukhina,b,c Josefina Perles,d Lai Feng,e,f Takeshi Akasaka,*e,ghi Tomas Torres*x,b,c and Dirk M. Guldi*a

Two subphthalocyanine–C₆₀ conjugates have been prepared by means of the 1,3-dipolar cycloaddition reaction of (perfluoro) or hexa(pentylsulfonyl) electron deficient subphthalocyanulines to C₆₀. Comprehensive assays regarding the electronic features — in the ground and excited state — of the resulting conjugates revealed energy and electron transfer processes upon photoexcitation. Most important is the unambiguous evidence — in terms of time-resolved spectroscopy — of an ultrafast oxidative electron transfer evolving from C₆₀ to the photoexcited subphthalocyanulines. This is, to the best of our knowledge, the first case of an intramolecular oxidation of C₆₀ within electron donor–acceptor conjugates by means of only photoexcitation.

porphyrinoid chromophores featuring unique panchromatic absorptive, redox, and electrical properties is of crucial importance.

To this end, fullerences – as one of the most explored class of molecular materials – have generated an enormous interest in the field of nanoelectronics. On one hand, their rigid aromatic structure evokes low reorganization energies in electron transfer reactions and, on the other hand, their extended π-conjugation affords efficient charge stabilization.3 Unique architectural flexibility and chemical versatility of fullerences allow the fine-tuning of their properties by, for example, covalent modification of their exterior4 or incorporation of molecular guests into their interior.5 In light of the aforementioned, a myriad of electron donor–acceptor conjugates/hybrids have been designed featuring porphyrins,6 (sub)phthalocyanulines,7,8 perylene-diimides,9 and other chromophores as light-harvesting electron donors and fullerences as electron acceptors. Notably, reports in which fullerences are employed as electron donors remain, however, scarce.

Oxidation of fullerences requires harsh conditions and, thus, occurs only by treating them with, for example, strong oxidizing agents10 or upon photo- or electron-induced ionization in the presence of electron transferring photosensitizers.11 The first example of intramolecular electron-transfer oxidation of fullerences was reported by Fukuzumi et al. in the presence of p-benzoquinone and radical ion pair stabilizing scandium triflate Sc(OTf)₃.12 Only a single case of intramolecular oxidation of fullerences has been reported to date. In a fullerence–trinitrofluorenone conjugate, photoexcitation and the presence of Sc(OTf)₃ leads, indeed, to the formation of oxidized fullerences, whereas only the triplet excited state of the fullerence was observed in the absence of Sc(OTf)₃.13 No successful
intramolecular oxidation of fullerenes by means of simple photoexcitation has ever been reported.

In stark contrast, recent studies have demonstrated that endohedral metallofullerenes function as electron donors in electron donor-acceptor conjugates. Thus, linking fullerenes to strong electron acceptors, which provide the means of charge stabilization within their extended \( \pi \)-system, could assist in shedding light onto the rare electron donating features even of empty fullerenes. This could open up new ways of using fullerenes in novel photoelectrochemical water splitting devices, which operate under oxidative conditions, novel solar cells, which feature larger open circuit voltages, or towards optoelectronic devices integrating novel n-type semiconductors.

In this connection, 14 \( \pi \)-electron aromatic subphthalocyanines (SubPc) represent ideal candidates with an extended conjugation and a cone-shaped geometry. Owing to their intense absorption in the visible region of the solar spectrum – complementary to that of fullerenes – and high optical and thermal stabilities, they have evolved as perfect light-harvesting building blocks. Low reorganization energies in electron transfer reactions and low reduction potentials of SubPcs, which bear electron-withdrawing substituents, render them ideal electron acceptors when linked to fullerenes.

Herein, we report on the synthesis as well as the electrochemical and photophysical features of conjugates 1a,b (Scheme 1) that is, \( C_{60} \) and electron accepting SubPcs with electron-withdrawing fluorine or pentyloxysulfanyl substituents on their periphery. This is meant to assist in realizing a game changer for advanced solar energy conversion schemes, namely photoinduced oxidation of empty fullerenes under small reorganization energies.

Results and discussion

Electron-deficient chloro-(alkylsulfanyl) SubPcs are known as the least reactive SubPcs in terms of axial substitution of the chlorine atoms. In fact, examples of conjugates involving (alkylsulfanyl) SubPcs are limited to only a single case, namely a hexa(octylsulfonyl) SubPc–ferrocene conjugate reported by our group. In the present work, (perfluoro) and hexa(pentyloxysulfanyl) derivatives of (formylphenoxy) substituted SubPc precursors 2a,b and (phenoxy) SubPc references 2c,d were successfully synthesized through triflate-substituted intermediates, generated by treating (chloro) SubPcs 3a,b with AgOTf (Scheme 1) according to a recently developed procedure. Corresponding triflate-containing SubPc intermediates were further subjected to in situ substitution of the triflyl group by 3-hydroxybenzaldehyde or phenol in toluene at 100 °C affording the desired (formylphenoxy) (2a,b) or (phenoxy) (2c,d) substituted SubPcs in good yields. Transformation of (chloro)-hexa(pentyloxysulfanyl) SubPcs 3b into the corresponding phenoxy-derivative 2d enabled its single crystal and, in turn, structural analysis by means of X-ray diffraction (Fig. S45 and S46 in the ESI†).

Conjugates 1a,b were obtained by the 1,3-dipolar cycloaddition of azomethine ylides generated in situ in a reaction of (3-formylphenoxy) SubPcs 2a,b and \( N \)-methylglycine, to \( C_{60} \). TLC and HPLC analyses (Fig. S1 and S2†) inform that the reaction time for forming 1b is only 2 h. This reflects the perturbation stemming from the six electron deficient substituents on the electronic properties of 2b and its corresponding intermediates. In contrast, perfluorinated SubPc 2a reacts in 8 h and reveals, as such, a reactivity similar to that seen for electron-donating SubPcs in Prato protocols. Subsequent column chromatography on silica gel, followed by size exclusion separation on Bio-Beads, yielded pure 1a in 63% and moderate amounts, namely around 6% of bis-adducts. Efficient purification of 1b was, nevertheless, hampered by its degradation on silica gel and formation of decomposition products with similar polarities. Performing column chromatography on cyanopropyldichlorosilyl-modified silica gel as a solid support allowed the successful separation of the conjugate in somewhat lower yields of 24% followed by the final separation of bis-adducts by GPC.

All compounds were fully characterized by \( ^1 H \) and \( ^{13} C \) NMR, MALDI-MS, UV-vis, and FTIR spectroscopy (Fig. S3–S38†). 1a and 1b are obtained as a mixture of two diastereoisomers as is known for Prato addition reactions to \( C_{60} \).

Being inseparable by conventional column chromatography on silica gel or by HPLC on a Buckyprep column (Fig. S1 and S2†), their presence is clearly manifested by the broadening in the \( ^1 H \) NMR spectra (Fig. S3 and S9†). The proton signals of the phenyl spacer in 1a and 1b are assigned on the basis of the coupling patterns observed in \( ^1 H \) COSY NMR (Fig. S4 and S10†). We conclude from the \( ^1 H \) NMR spectra that the proton signals of the phenoxy spacer remain unaffected by the nature of the SubPc substituents and in both cases appear shifted upfield into the range from 5.4 to 6.9 ppm as a consequence of the SubPc ring current.

The electrochemical properties of 1a,b were examined and compared with the redox characteristics of their references, that is, 2c,d and \( N \)-methylfulleropyrrolidinone 4, by CV (Fig. 1) and DPV (Fig. S39†) in o-DCB containing 0.05 M \(( n-Bu)_4NPF_6\). The corresponding data are summarized in Table 1. Overall, 1a and 1b exhibit up to seven reductions and two oxidations in the cathodic and the anodic directions, respectively, which were assigned upon comparison with SubPcs 2c,d and \( N \)-methylfulleropyrrolidinone 4.

Scheme 1 Synthesis of SubPc–\( C_{60} \) conjugates 1a,b. Conditions: (a) AgOTf, 3-hydroxybenzaldehyde/phenol, toluene; (b) \( N \)-methylglycine, \( C_{60} \), toluene, reflux.
For example, (perfluoro) SubPc 2c reveals three reversible reductions at −1.09, −1.71, and −2.30 and one reversible oxidation at 1.02 V. Similarly, four reversible reductions at −0.81, −1.34, −1.73, and −2.19 V and one irreversible oxidation at 1.18 V are noted for hexa(pentylsulfonyl) SubPc 2d. With the latter in hand, the reductions of SubPc-C$_{60}$ conjugate 1a evolving at −1.18, −1.79, and −2.21 V are ascribed to SubPc centered processes, whereas those at −1.06 and −1.57 V relate to processes involving C$_{60}$ (Table 1). Oxidations at +0.99 and +1.08 V are C$_{60}$ and SubPc centered, respectively. In 1b, the presence of two strong acceptors, namely hexa(pentylsulfonyl) SubPc and C$_{60}$, leads to a coalescence and irreversibility of the reductions followed by two oxidations. As such, seven reductions at −0.91, −1.10, −1.26, −1.42, −1.62, −1.81, and −2.17 V plus two oxidations at +1.04 and +1.12 V evolve (Table 1). Those at +1.04, −1.10, −1.42, and −1.81 V are C$_{60}$ centered, while those at +1.12, −0.91, −1.26, −1.62, and −2.17 V correlate with either the oxidation or the reduction of hexa(pentylsulfonyl) SubPc.

From Fig. 1 and Table 1, we conclude that hexa(pentylsulfonyl) SubPc 2d is more easily reduced and more difficult to be oxidized than (perfluoro) SubPc 2c. In other words, the alkysulfonyl groups impose stronger electron-withdrawing character than the fluorine atoms. A corroborating trend is found in 1a and 1b, 1b is more easily reduced and more difficult to be oxidized than 1a, where the differences in reduction and oxidation are 150 and 50 mV, respectively. On the other hand, the corresponding first reductions in 1a and 1b are considerably 30 and 100 mV shifted if compared to those of the reference SubPcs 2c and 2d, suggesting stronger ground-state interactions between the electron donor and the electron acceptor subunits in 1b. The HOMO–LUMO gaps were calculated using the data (Table 1). The estimated values are found to be in good agreement with the SubPc ground state absorption features. Thus, 1b containing hexa(pentylsulfonyl) SubPc reveals strong absorption at 582 nm, which is 7 nm red-shifted in comparison with the absorption of perfluorinated 1a at 575 nm (Fig. S7 and S13†). The precursors (2a,b) and the reference SubPcs (2c,d) follow the same trend. As such, the absorption spectra of 1a and 1b are best described as the superimposition of the absorption characteristics of the single components, namely SubPc 2c and d, on one hand, and N-methylfulleropyrrolidine 4, on the other hand.

In line with what was observed in steady state absorptions, the nature of the peripheral substituents influences the singlet excited state properties of 1a and 1b (Fig. 2). In reference experiments, a solvent independent fluorescence quantum yield of 0.17 is determined for 2c. In contrast, 2d exhibits solvent dependent quantum yields of 0.15, 0.07 and 0.06 in toluene, THF and benzonitrile, respectively. In reference to the aforementioned, fluorescence assays with 1a and 1b give rise to strong quenching of the SubPc fluorescence with values from 0.0007 to 0.0012.

![Fig. 1](image1.png) Cyclic voltammograms of 1a, 2c, 2d and 4 (10$^{-3}$ M) and 1b (10$^{-4}$ M) measured at 100 mV s$^{-1}$ in o-dichlorobenzene with 0.05 M (n-Bu$_4$)NPF$_6$ as supporting electrolyte, glassy carbon as working electrode, Pt wire as counter electrode, and Ag wire as reference electrode.

![Fig. 2](image2.png) Absorption (solid lines) and fluorescence spectra (dashed lines) of SubPc 2c (black lines) and conjugate 1a (red lines) in toluene.

<table>
<thead>
<tr>
<th>Subpces</th>
<th>E$_{p,red}$</th>
<th>E$_{p,ox}$</th>
<th>E$_{p,ox}$</th>
<th>E$_{p,ox}$</th>
<th>E$_{p,ox}$</th>
<th>HOMO–LUMO gap</th>
</tr>
</thead>
<tbody>
<tr>
<td>1a SubPc centered</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>C$_{60}$ centered</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>1b SubPc centered</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>C$_{60}$ centered</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>1c</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>1d</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>2a</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>2b</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>2c</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
<tr>
<td>2d</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
<td>2.05</td>
</tr>
</tbody>
</table>
To assign the differential absorption changes – *vide infra* – spectroelectrochemical reductions of 2c or 2d to form the corresponding \([\text{SubPc}(F_{12})]^{+}\) or \([\text{SubPc}((\text{SO}_2\text{C}_5\text{H}_{11})_6)]^{+}\) radical anions were deemed important. On one hand, reduction of 2c results in two maxima at 455 and 655 nm, two shoulders at around 475 and 610 nm, and a minimum at 570 nm. On the other hand, upon reduction of 2d spectral characteristics with maxima at 480, 545, 620, and 735 nm complemented by minima at 534 and 581 nm evolve (Fig. S44). Notably, pulse radiolysis reductions with 2c or 2d in deaerated toluene–2-propanol–acetone mixtures (8:1:1 v/v) result in quantitatively similar spectra with characteristic fingerprints at 610 and 620 nm, respectively. For establishing the signature of \([\text{C}_{60}]^{+}\) we turned to time-resolved investigations. From, for example, pulse radiolysis experiments with N-methyl-3,4-fulleropyrrolidine 4 we derive characteristic maxima for \([\text{C}_{60}]^{+}\) in the near infrared at 860 and 950 nm. As a complement, we probed the scission-ion promoted photoinduced oxidation of 4 by \(p\)-chloranil in deaerated benzonitrile. Here, a broad signature in the near infrared with maxima at 880, 960, and 1080 nm corroborate the pulse radiolysis experiments with 4. Important is the determination of the extinction coefficient for the 880 nm maximum of approximately 3000 M\(^{-1}\) cm\(^{-1}\).

Insights into the excited state characteristics of 4, 2c, 2d, 1a, and 1b came from transient absorption measurements following femtosecond and nanosecond excitation. For N-methylfulleropyrrolidine, differential absorption changes evolve immediately after the 887 nm laser excitation, which are characterized by maxima at 500 and 900 nm. The \([\text{C}_{60}]^{+}\) singlet excited state (1.8 eV) decays with 1.4 ± 0.1 ns to afford the energetically lower lying triplet excited state (1.5 eV), whose characteristics are a transient maximum at 700 nm and a transient lifetime of up to 20 \(\mu\)s in the absence of molecular oxygen. 2c reveals, upon excitation at 530 nm, differential absorption changes, which include transient maxima at 440 and 600 nm as well as transient minima at 514, 575, and 635 nm (Fig. S40†). In addition, a broad near infrared feature spans from 650 to 1200 nm, which peaks around 710 nm. These features relate to the SubPc singlet excited state (2.16 eV) and transform with 1.9 ± 0.1 ns into the corresponding triplet excited state (1.4 eV).Transient absorption spectra of the latter maximize at 470 and 610 nm and minimize at 532 and 570 nm. They all give rise to the same triplet excited state lifetime of 39 ± 2 \(\mu\)s. For 2d, commencing with the conclusion of the 530 nm excitation, differential absorption changes develop in the form of transient maxima at 424, 474, 623, and 660 nm, a broad tail extending far into the near infrared, as well as transient minima at 533 and 583 nm (Fig. S41†). These SubPc singlet excited state (2.12 eV) related transient absorption features undergo intersystem crossing to the corresponding triplet excited state (1.4 eV), which exhibits a broad transient in the visible part of the spectrum that maximizes at 470 and 620 nm and minimizes at 533 and 583 nm. Owing to the presence of the sulfur, singlet excited state lifetimes of 1220 ± 20, 420 ± 10, and 415 ± 10 ps evolve in toluene, THF, and benzonitrile, respectively. In contrast, the triplet excited state lifetime is 53 ± 2 \(\mu\)s.

530 nm excitation of 1a results in the exclusive formation of the SubPc singlet excited state (Fig. S42†). In particular, transient maxima at 450, 600, and 720 nm as well as transient minima at 515, 575, and 635 nm are formed instantaneously and decay in the presence of \([\text{C}_{60}]^{+}\) rapidly with about 1.5 ± 0.3 ps in toluene, THF, and benzonitrile. As the SubPc singlet excited state decays a broad near-infrared transient, which maximizes at 910 nm, grows in, and, whose features are attributed to the \([\text{C}_{60}]^{+}\) singlet excited state. Interestingly, we did not find the characteristic \([\text{C}_{60}]^{+}\) triplet feature at the end of the \([\text{C}_{60}]^{+}\) excited state deactivation. On the contrary, maxima at 470 and 615 nm as well as a minimum at 575 nm are noted. Earlier we have established such features as reliable attributes of the SubPc triplet excited state. From this we infer that the \([\text{C}_{60}]^{+}\) triplet excited state (1.5 eV), once formed, undergoes a thermodynamically allowed transfer of triplet excited state energy to SubPc (1.4 eV).

The kinetics at the 470 and 615 nm maxima, which allowed us to follow the generation of the SubPc triplet excited state, further furnish the following kinetic assignment: the rate-determining step in the SubPc triplet excited state formation is the \([\text{C}_{60}]^{+}\) centered intersystem crossing. A global analysis discloses kinetics that are very similar (1.6 ± 0.1 ns) to the inherent intersystem crossing dynamics seen for \([\text{C}_{60}]^{+}\). When turning to excitation of 1b at 530 nm (Fig. S43†), its singlet excited state, that is, maxima at 427, 474, 623, and 660 nm, and minima at 535 and 585 nm, decay in the presence of \([\text{C}_{60}]^{+}\) rapidly within 1.5 ± 0.3 ps to form accordingly the \([\text{C}_{60}]^{+}\) singlet excited state with its 910 nm maximum. Like for 1a, we did not find the characteristic \([\text{C}_{60}]^{+}\) triplet feature. Instead, maxima at 470 and 625 nm as well as minima at 535 and 585 nm of the SubPc triplet excited state were concluded. Again, the triplet excited state energy of SubPc (1.4 eV) evolves from a thermodynamically allowed transfer of triplet excited state energy. The kinetics document that the rate-determining step is the \([\text{C}_{60}]^{+}\) centered intersystem crossing (1.6 ± 0.1 ns).\(^{21}\)

Exciting 1a at 320 or 387 nm populates higher lying singlet excited states \((S_2: 3.4 ± 0.1 \text{ eV})\) of SubPc directly – the light partitioning at, for example, 320 nm is 1 : 1 (Fig. 3). The latter are discernible with maxima at 430, 600, and 995 and minima at 515 and 570 nm, which transform within 7 ± 1 ps to the lowest-energy singlet excited state in the absence of \([\text{C}_{60}]^{+}\). In contrast to that, in 1a the lifetime is shortened to 4.1 ± 0.2 ps (toluene), 3.5 ± 0.2 ps (THF), and 2.6 ± 0.5 ps (benzonitrile) by the presence of \([\text{C}_{60}]^{+}\). The product of an instantaneous transformation includes maxima at 590, 840, 910, and 1020 nm and a minimum at 575 nm. These are in sound agreement with those noted upon pulse radiolylotical generation of \([\text{C}_{60}]^{+}\) – *vide supra*. In the visible, a transient maximum at 595 nm and a transient minimum at 578 nm are clear fingerprints of \([\text{SubPc}(F_{12})]^{+}\). In other words, the close spectral resemblance between radiolylotical and photophysical experiments prompts to the unprecedented formation of the \([\text{C}_{60}]^{+}\)–\([\text{SubPc}(F_{12})]^{+}\) \((2.17 \text{ eV})\) radical ion pair state, as a product of charge separation (Fig. 4). Both fingerprints served as reliable probes to determine the lifetime of the metastable \([\text{C}_{60}]^{+}\)–\([\text{SubPc}(F_{12})]^{+}\) radical ion pair state. Charge recombination proceeds with time constants of 31 ± 5

4144 | Chem. Sci., 2015, 6, 4141–4147

This journal is © The Royal Society of Chemistry 2015
ps (toluene), 35 ± 5 ps (THF), and 38 ± 10 ps (benzonitrile) to populate the SubPc triplet excited state as the only discernable component. The singlet excited state of C$_{60}$ (1.8 eV), which is partly excited at excitation wavelengths like 320 and 387 nm, is, however, insufficient in terms of thermodynamics to drive a charge transfer. Instead, the C$_{60}$ singlet excited state related transient absorptions with maxima at 500 and 900 nm transform with a time constant of 1.6 ± 0.1 ns into the SubPc triplet excited state via triplet excited state energy transfer, similar to the 530 nm excitation experiments. However, the higher lying singlet excited SubPc states (S$_2$: 3.3 ± 0.2 eV) are able to drive an ultrafast electron transfer evolving from C$_{60}$ to SubPc. As a matter of fact, maxima at 481, 623, 655, and 1065 nm and minima at 532 and 585 nm decay fast in toluene, THF, and benzonitrile with lifetimes of around 0.6 ± 0.2, 1.4 ± 0.5, and 1.7 ± 0.5 ps, respectively. Simultaneously with the latter decay, new transitions develop in the visible and the near-infrared regions. Importantly, the new transients do not match the signature of the SubPc triplet excited state. Instead, maxima at 480, 550, 610, 840, 910, and 1025 nm as well as a

Fig. 4 Energy level diagram of 1a reflecting the different pathways of energy and electron transfer.
minimum at 590 nm are discernable. Upon spectral comparison – *vide supra* – we ascribe the transitions in the visible to (SubPc(SO2C5H11)6)–, while those in the near-infrared correspond to (C60)–. In light of the aforementioned, we conclude an electron transfer from C60 to, for example, the SubPc singlet excited state to yield (C60)–→(SubPc(SO2C5H11)6)– (1.95 eV). The (C60)–→(SubPc(SO2C5H11)6)– radical ion pair state is short lived and decays with lifetimes of 16 ± 5, 21 ± 5, and 26 ± 5 ps in toluene, THF, and benzonitrile, respectively. At the end of this decay, the SubPc triplet excited state with maxima at 470 and 625 nm as well as minima at 535 and 585 nm emerges as the product of charge recombination.2

Conclusions

To sum up, we have documented the comprehensive investigation regarding the covalent attachment of two different light harvesting/electron accepting SubPcs to C60 by means of the Prato reaction complemented by several steady-state and time-resolved assays. Most importantly, our photophysical assays corroborate unambiguously the ultrafast charge separation (in the range from 0.6 to 4.1 ps) between SubPcs and C60 upon photoexcitation at either 320 or 387 nm to form the one-electron oxidized form of C60 and the one-electron reduced forms of the two different SubPcs. This is, to the best of our knowledge, the first case of an *intramolecular* oxidation of C60 as part of an electron donor–acceptor conjugate for advanced solar energy conversion schemes by means of only photoexcitation. The charge separated state lifetime is at this state rather short lived with values of up to 38 ps. Currently, we are directing our efforts to extend the radical ion pair state lifetime by, for example, changing the energetics/thermodynamics of charge recombination.

Acknowledgements

Financial support from the MICINN and MEC, Spain (CTQ-2011-24187/BQU and PRI-PIBUS-2011-1128) is acknowledged. D.M.G. appreciates funding by the DFG (GU 517/14-1) and “Solar Technologies Go Hybrid” – an initiative of the Bavarian State Ministry for Science, Research, and Art.

Notes and references


Irreversibility is likely based on electron density rearrangement chemical reactions.

In this context, it is reassuring that the transients seen at the end of the femtosecond experiments match those at the beginning of the nanosecond experiment. Moreover, maxima at 470 and 610 nm, minima at 532 and 570 nm, and an excited state lifetime of 36 μs in the absence of oxygen perfectly agree with what was found for the SubPc triplet excited state of 2c.

In the absence of oxygen, the SubPc triplet excited state lifetime is 20 ± 5 μs and, as such, agrees with what was found for 2d.

Singlet oxygen quantum yields of 1b were found to be as high as 0.49 (toluene), 0.24 (THF), and 0.65 (benzonitrile) and support the assignment that the triplet excited state evolves as the product of charge recombination. Please note that the singlet oxygen yields in 2d are 0.39 (toluene), 0.09 (THF), and 0.26 (benzonitrile).