

Article

Improving the Flame Retardance of Polyisocyanurate Foams by Dibenzo[d,f][1,3,2]dioxaphosphepine 6-Oxide-Containing Additives

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Abstract: A series of new flame retardants (FR) based on dibenzo[d,f][1,3,2]dioxaphosphepine 6-oxide (BPPO) incorporating acrylates and benzoquinone were developed previously. In this study, we examine the fire behavior of the new flame retardants in polyisocyanurate (PIR) foams. The foam characteristics, thermal decomposition, and fire behavior are investigated. The fire properties of the foams containing BPPO-based derivatives were found to depend on the chemical structure of the substituents. We also compare our results to state-of-the-art non-halogenated FR such as triphenylphosphate and chemically similar phosphinate, i.e. 9,10-dihydro-9-oxa-10-phosphaphenanthrene-10-oxide (DOPO), based derivatives to discuss the role of the phosphorus oxidation state.

Keywords: flame retardant; dibenzo[d,f][1,3,2]dioxaphosphepine 6-oxide; DOPO; phospha-Michael addition; polyisocyanurate

1. Introduction

Polyurethanes that are synthesized by polyaddition of polyisocyanates and polyols to generate a class of materials with an extremely adaptable property profiles [1]. The molar excess of the polyisocyanate results in crosslinked polyisocyanurate (PIR) structures. The isocyanurate rings increase the stability of the material as compared to the corresponding polyurethanes [2–5]. Furthermore, crosslinked PIR foams do not melt. Both crosslinking and isocyanurate rings result in enhanced char formation after burning. Therefore, these properties as well as their closed cell structure, excellent strength at low density, and low thermal conductivity PIR foams are particularly well-suited as insulation materials [1] e.g., as insulation panels in constructions. However, a need for reduction of energy costs requires more efficient insulation, and recent incidents underline the importance of efficient fire protection in building and construction [1,6,7]. Of particular importance is the formation of insulating char on the surface of buildings to protect the structures from excessive heat.

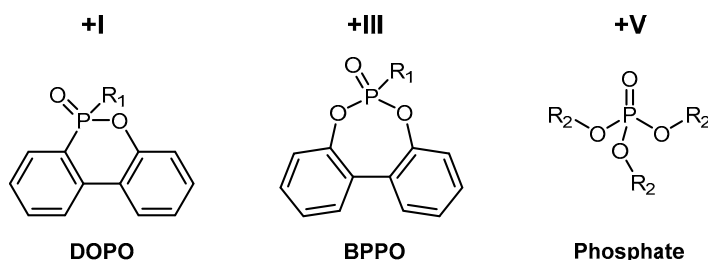
For PIR, additive tris(2-chloro isopropyl)phosphate (TCPP) is the state-of-the-art flame retardant (FR) [8,9]. It combines both halogen and phosphate in one molecule and it leads to a good balance in gas phase versus solid phase FR action in the resulting foams. TCPP also lowers the viscosity of the formulation for improved processing and plasticizes the foam [10]. However, there is interest in halogen-free FRs to comply with certain eco-labels, for instance the label “pure life” [11]. Phosphorus (P)-containing molecules are considered to be substitutes for halogenated FRs [9,12,13].

Phosphorus offers a variety of structures with different oxidation states of the phosphorus atom [14]. Generally speaking, the mode of action, i.e., gas phase versus solid phase action, is dominated by two aspects for phosphorous based flame retardants [2,15–18]. First, the oxidation state of the phosphorus atom in the compounds plays a decisive role. If the phosphorus atom has a high oxidation state a thermally stable char is more likely to be formed and the emission of P-containing gases (gas phase mechanism) is low. If the phosphorus oxidation state is low, P-containing gases are more likely to be released. Gas phase action is strongest for phosphine oxides and low for [16]. Secondly, the volatility of the FR impacts the gas phase activity [19]. Balancing gas phase and condensed phase activity in polymer materials is important to tune FR performance and comply with safety standards [17,18,20].

Isobutyl bis(hydroxymethyl) phosphine oxide with an oxidation state of P of $-I$ is an example of an FR agent with high gas phase activity [17]. Phosphinates, e.g., aluminum diethyl phosphinate [18,21], 9,10-dihydro-9-oxy-10-phosphaphenanthrene-10-oxide (DOPO), and its derivatives [13,22,23] have an oxidation state of $+I$ and they are well-known FR additives for many polymeric materials. DOPO derivatives (either as free, non-reactive additive, or incorporated into a polymer backbone) mostly introduce a combination of gas phase/condensed phase action and intumescence [24–26]. Derivative incorporating secondary amines yielded phosphonamidates that were applied as FR in polyurethanes [27]. Phosphonates with intermediate oxidation state of $+III$, e.g., diethyl ethylphosphonate (DEEP) and dimethyl phenylphosphonate (DMPP) have been established as halogen-free FRs for polyurethanes, polyesters, and polyamides [28]. In all cases, the P-containing compounds were applied as non-reactive FR. DOPO [29], phosphates [20,30], and phosphinic oxide compounds [17,31] can also be used to prepare P-containing diols that are employed as reactive FRs in polyurethane formulations.

Here, we explore the efficiency of dibenzo[d, f][1, 3, 2] dioxaphosphepine 6-oxide (BPPO) compounds as FR in PIR foams. The BPPO compounds are phosphonates with an oxidation state of $+III$ and are compared to the DOPO compounds with an oxidation state of $+I$. Triphenyl phosphate (TPP), a state-of-the-art non halogenated FR with an oxidation state of $+V$, served as benchmark. The structural motives are presented in Figure 1. Triethyl phosphate (TEP), a common additive in PIR foams, was also part of most formulations. TEP acts as FR but also plastics the material, improving mechanical properties of the PIR foams. This study aims to investigate differences in phosphonate, phosphinate and phosphate (TPP) based FR in technically relevant PIR formulations.

Phosphorus oxidation state:


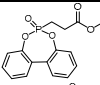

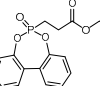

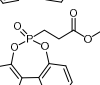

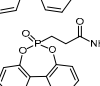

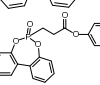

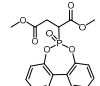

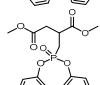
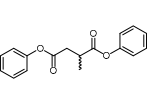
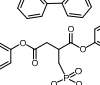

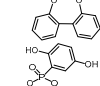

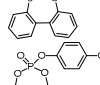

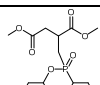

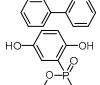


R_1 = see Table 2, $R_2 = CH_2CH_3$ = Triethyl phosphate, $R_2 = C_6H_5$ = Triphenyl phosphate

Figure 1. Chemical structures of 9,10-dihydro-9-oxy-10-phosphaphenanthrene-10-oxide (DOPO), dibenzo[d,f][1,3,2] dioxaphosphepine 6-oxide (BPPO), and the described phosphates with the assigned oxidation state of the phosphorus atom.

The synthesis of the novel BPPO compounds that are listed in Table 1 has been previously reported by our group [32]. We found that the synthesis of BPPO proceeds under significantly milder conditions than those reported for DOPO [33].

Table 1. BPPO derivatives studied as flame retardant (FR) additives in polyurethane foams.

Entry	BPPO Added to...	R ₁	Chemical Structure	Abbreviation
1	Methyl acrylate			MA-BPPO
2	Ethyl acrylate			EA-BPPO
3	Tert. Butyl acrylate			tBuA-BPPO
4	Acrylamide			AM-BPPO
5	Phenyl acrylate			PA-BPPO
6	Dimethyl fumarate			SU-BPPO
7	Dimethyl itaconate			DMI-BPPO
8	Diphenyl fumarate			DPF-BPPO
9	p-Benzoquinone			HQ-BPPO
10	p-Benzoquinone			HP-BPPO
DOPO added to...				
11	Dimethyl itaconate			DMI-DOPO
12	p-Benzoquinone			HQ-DOPO

The addition of BPPO to unsaturated compounds via a Phospha-Michael addition [23,34] more readily proceeds and the products are easier to purify. The FR efficiency of the new BPPO compounds are compared in rigid PIR foams with DOPO compounds and with benchmark foam with TEP and TPP. The formulations were kept constant according to Table 2, and only the amount of FR was varied to maintain a constant P-content to ensure the comparability of the foams. The stepwise variation of the chemical structure of the FR provides the opportunity to assign the occurring effects to structural features.

Table 2. Formulation of the polyisocyanurate (PIR) foams studied.

Ingredient	Amount [g]
Stabilizer (TEGOSTAB B 8421)	4.0
Emulsifier (Emulsogen TS100)	2.0
Triethyl phosphate (TEP)	5.0
PEG 400	16.0
Catalyst (KAc 25 wt.% in diethylene glycol)	2.5
Polyester polyol (PEP50 AD)	53.0
Blowing Agent (pentane)	15.0
Flame Retardant ^a	x
Polyisocyanate (DESMODUR 44V70L)	151.8

^a Amount depends on the P-content of the FR according to Table 1 itself and the P-content desired (1 wt.%) in the foam.

2. Materials and Methods

2.1. Chemicals

Pentane ($\geq 99\%$, Sigma-Aldrich, Darmstadt, Germany); triphenyl phosphate (TPP, $>99.0\%$, TCI, Eschborn, Germany), triethyl phosphate (TEP, $>99.0\%$, TCI, Eschborn, Germany); PEG 400 (Sigma-Aldrich, Darmstadt, Germany), $n\text{OH} = 280\text{ mg KOH/g}$; potassium acetate (KAc, $\geq 99.0\%$, Sigma-Aldrich, Darmstadt, Germany); diethylene glycol (DEG, $>99.0\%$, Sigma-Aldrich, Darmstadt, Germany); phthalate/DEG polyester polyol PEP50AD (Covestro Deutschland AG, Dormagen, Germany), $n\text{OH} = 240\text{ mg KOH/g}$; polymeric isocyanate DESMODUR 44V70L, $\text{NCO}\% = 30.9$ (Covestro Deutschland AG, Dormagen, Germany); Emulsogen TS100 (Clariant, Pratteln, Switzerland), polyether modified polysiloxane TEGOSTAB B 8421, $n\text{OH} = 57\text{ mg KOH/g}$ (Evonik Industries, Darmstadt, Germany); p-tolyl isocyanate (99% , Sigma-Aldrich, Darmstadt, Germany), 1,4-dioxane (99.8% , Sigma-Aldrich, Darmstadt, Germany) were used as received. For PIR foam preparation, a solution of 25 wt% potassium acetate in DEG ($n\text{OH}$: 793 mg KOH/g) was used as a catalyst. The synthesis of the BPPO-containing FRs according to Table 1 has been recently described [32]. The DOPO-derivative of dimethyl itaconate (DMI-DOPO) was synthesized according to Pospiech et al. [35].

2.2. Foam Preparation

The preparation of the PIR foams was performed in two steps. First, a mechanical stirrer mixed the polyester polyol, stabilizers, surfactants, catalyst, blowing agent, and FR at 2000 rpm. Subsequently, the required amount of isocyanate was added. After mixing for a few seconds, the mixture was poured into an open mold to prepare the foam. During the foaming process, three distinct times (cream time, setting time, and rise time) were noted all starting from the mixing of polyol formulation and isocyanate. Cream time is defined as the point at which the mixture turns creamy and starts to expand. The setting time, which is also called gel time or fiber time, reflects the time when solid fibers can be drawn from the expanding foam. The rise time indicates when the foam expansion is completed. After the foam expanded, it was removed from the mold and kept at room temperature for 24 h before the specimens were cut with a band-saw.

Table 2 summarizes the general composition of the foams. The complete formulations are listed in Table S1 in the Supplementary Materials.

2.3. Characterization of Foam Properties

The density, pore sizes, and cell integrity of the foams were characterized. The samples used for cone calorimetry ($10\text{ cm} \times 10\text{ cm} \times 5\text{ cm}$) were employed to calculate the density and their volume and weight was determined. Pore sizes were determined from light microscopic images. The

microscope used was an Axio Imager (ZEISS) that was equipped with AxioCam 305 color camera (ZEISS, Oberkochen, Germany).

The scanning electron microscopy (SEM) images were taken by a Gemini Ultra plus SEM (ZEISS, Oberkochen, Germany). The water absorption was determined as a parameter for the cell integrity. Foam cubes of 4 cm × 4 cm × 4 cm were completely immersed in boiling water for 90 min. The mass difference ($m_w - m_d$) was determined (m_w describes the mass of the wet foam, m_d mass of the dry foam). With the density of water (1 g·cm⁻³), the mass difference was divided by the sample volume (a^3) and normalized to 100%. With the following Equation (1), the water absorption WA_V (volume of absorbed water over foam volume) was obtained.

$$WA_V = \frac{m_w - m_d}{a^3 \times \rho} \times 100\% \quad (1)$$

2.3.1. Mechanical Properties of the Foams

The compressive strength of the foams was investigated by a TIRAtest 2300 instrument (TIRA GmbH, Schmalkalden, Germany). The foams were tested according to ISO 604 while using a 100 kN force sensor. A traverse was used for the position sensor. An initial force of 2 N was applied and the test was conducted with a velocity of 500 mm·min⁻¹. The test specimen cubes with edge lengths of 5 cm were used. The compression curves were perpendicularly measured to the foam rise direction.

2.3.2. ATR-FTIR Analysis

The infrared spectra were measured with a Vertex 80v spectrophotometer (Bruker, Rheinstetten, Germany) equipped with a golden gate diamond ATR unit (SPECAC) in the wavenumber range of 4000–600 cm⁻¹ with 100 scans per measurement. A MCT (Mercury cadmium telluride, Bruker, Rheinstetten, Germany) detector was used with a resolution of 4 cm⁻¹.

2.3.3. Quantitative Phosphorus Content

The quantitative P-content was determined by Mikroanalytisches Labor Kolbe (Mülheim a.d. Ruhr, Germany).

2.3.4. Thermal Decomposition

The thermal behavior was assessed by thermogravimetric analysis (TGA) using a TGA Q500 (TA Instruments, New Castle, UK) in nitrogen atmosphere (60 mL·min⁻¹) from 25 to 800 °C at a scan rate of 10 K·min⁻¹ with a sample weight of 5 mg. The thermal decomposition products were analyzed with pyrolysis-GC/MS. These experiments were carried out with a GC 5890 (Agilent Technologies, Santa Clara, USA) coupled with a pyroprobe 2000 (CDS Instruments, Oxford, USA) under helium atmosphere and a flow rate of 1.0 mL·min⁻¹.

2.4. Fire behavior

2.4.1. Vertical Flame Spread

The vertical flame spread (VFS) test was carried out according to DIN 4102 (B2 classification). Thus, samples with dimensions of 20 cm × 10 cm × 1 cm were vertically hung in the test chamber (here, a test chamber for UL-94 test was used). A burner flame was applied for 15 s on the lower edge of the specimen. The height of the flame was measured. The test was passed when the height of the flame was lower than or equal to 15 cm.

2.4.2. Cone Calorimeter

The prepared PIR foams were cut into specimens with dimensions of 10 cm × 10 cm × 5 cm. The samples were conditioned for 48 h at 23 °C in a climate chamber with 50% relative humidity. They were tested in a cone calorimeter after 48 h (Fire Testing Technology, East Grinstead, UK) according

The foams were analyzed by ATR-FTIR spectroscopy. The FTIR spectra of the foams are shown in the Supporting Information (Figures S1 and S2). The spectrum of the pure EA-BPPO (Figure S3 in the Supplementary Materials) showed intense bands at 717 (P-C stretching), 925 (P-Ar), and 1242 cm^{-1} (P=O stretching) [36]. In the FTIR spectrum of foam EA-BPPO/TEP-1.0%P with 7 wt.% EA-BPPO (matching 1 wt.% phosphorus in the foam) the same bands appeared with weaker intensity. The stretching vibration bands for urethane at 1218 cm^{-1} and isocyanurate at 1405 cm^{-1} indicated the successful reaction of the isocyanate groups with aliphatic OH groups of the polyol [37]. A residue of non-reacted isocyanate groups was observed at 2273 cm^{-1} . This is due to the high NCO index of 3.2 that is necessary for generating isocyanurate rings [37].

In the case of AM-BPPO, HQ-BPPO, and HP-BPPO, a reaction between isocyanate and amide groups (AM-BPPO) and aromatic OH groups (HQ-BPPO, HP-BPPO), respectively, was assumed to occur during foam generation. These reactions could not be proven by FTIR. Therefore, model reactions between p-tolyl isocyanate and FR in stoichiometric amounts catalyzed by potassium acetate were performed in DMSO-d₆ at 80 °C (a temperature that can be detected during foam preparation inside the material). After one hour, ¹H NMR indicated a reaction of HP-BPPO with the isocyanate. A reaction between AM-BPPO and isocyanate to acetylurea could not be proven. Consequently, it was concluded that HP-BPPO and HQ-BPPO belong to the group of reactive FRs [17]. The stability of the BPPO ring against OH groups was examined with EA-BPPO in ethanol in the presence of potassium acetate at 80 °C for 1 h. EA-BPPO was found to be stable under these conditions, as the ¹H NMR only showed the educts and no degradation products. The spectrum of EA-BPPO was preserved (as compared to NMR spectrum in [32]). Therefore, it was expected that the BPPO ring is stable in the foaming process. The BPPO compounds were subjected to transesterification with diols (ethylene glycol, 1,4-butane diol) catalyzed by Ti(OBu)₄ to obtain reactive FRs (Table S2 and Table S3). A reaction could not be observed under various conditions, in most cases the BPPO ring opened. For SU-BPPO and DMI-BPPO, the synthesis of oligomers with different diols, as outlined before for DOPO derivatives [38], was also examined (Table S2). These experiments were not successful and they showed ring opening of the BPPO phosphonate ring.

3.2. Foam Properties

The foams were prepared under comparable reaction conditions (see experimental section). The foam morphology, density, water uptake, and compression strength of the various samples were analyzed to investigate the influence of the composition on the physical properties. Table 3 summarizes the results. The sample names contain the type of FR as well as the total P-content in wt.%.

The P-contents of the foams that are listed in Table 3 were calculated from the starting foam formulation taking the contribution of all phosphorus compounds into account. In the foams with EA-BPPO, MA-BPPO, AM-BPPO, and HQ-BPPO, the content of FR was varied to achieve P-contents between 1.0 to 1.5 wt.%. The quantitative P-contents were analyzed by Mikroanalytisches Labor Kolbe (Germany) for selected samples (Table S5). The contents found were slightly higher (1.1 wt.% for TPP/TEP-1.0%P and 1.3 wt.% for EA-BPPO/TEP-1.0%P) than the calculated equivalents (1.0 wt.%).

The control foams Ref-0%P and TEP-0.3%P had densities in the range of 36 $\text{kg}\cdot\text{m}^{-3}$. The addition of TPP resulted in a slight increase to 39 $\text{kg}\cdot\text{m}^{-3}$. The BPPO compounds added at a concentration of 1 wt.% phosphorus yielded foams in the desired density range of 38–42 $\text{kg}\cdot\text{m}^{-3}$. DMI-DOPO presented a comparable result. This density is typical for foams with steel facings that are used in cold store construction. Higher BPPO FR contents increased the density, as in the case of DMI-DOPO. Foams that were prepared with AM-BPPO and HQ-BPPO showed densities of around 50 $\text{kg}\cdot\text{m}^{-3}$. This is attributed to the fact that kinetic parameters of foam preparation (cream time, setting time, rise time) were significantly lower than for all other foams. The BPPO additives, except DMI-BPPO, were powders that were milled to comparable particle size (50–100 μm) prior to use. It was noted that all BPPO compounds, except AM-BPPO, SU-BPPO, and HQ-BPPO, were soluble in the formulation at elevated temperature.

The foam appearance that was observed in SEM was independent of the solubility of the FR. In all cases, homogeneous foams without visible particles (note the starting particle size of 50–100 μm) within the foam cells were obtained. Figure 3 shows SEM images of foams TEP-0.3%P (Figure 3a), TPP/TEP-1.0%P (Figure 3b), MA-BPPO/TEP-1.0%P (with soluble MA-BPPO, Figure 3c), and AM-BPPO/TEP-1.3%P (with insoluble AM-BPPO Figure 3d) as examples. The pore sizes of the foams according to optical microscopy varied between 0.14 and 0.20 mm. The differences between the foams were not significant.

Table 3. Physical properties of the studied foams (description of methods: see Experimental part).

Foam	P-content	Density	W_{Av}	Pore size	Cream Time	Setting Time	Rising Time
	[wt.%]	[$\text{kg}\cdot\text{m}^{-3}$]	[vol%]	[mm]	[s]	[s]	[s]
Ref-0%P	0.0	36	7	0.15	12	30	41
TEP-0.3%P	0.3	36	4	0.18	10	23	31
TPP-0.7%P	0.7	36	7	0.17	16	30	40
TPP/TEP-1.0%P	1.0	39	5	0.15	10	24	32
TPP/TEP-1.3%P	1.3	40	10	0.14	14	25	35
TPP/TEP-1.5%P	1.5	39	12	0.10	13	23	30
MA-BPPO/TEP-1.0%P	1.0	38	4	0.17	10	35	45
MA-BPPO/TEP-1.2%P	1.2	39	14	0.15	14	38	48
MA-BPPO/TEP-1.4%P	1.4	37	9	0.17	13	48	58
EA-BPPO-0.7%P	0.7	36	7	0.29	16	40	45
EA-BPPO/TEP-1.0%P	1.0	37	6	0.19	9	26	33
EA-BPPO/TEP-1.3%P	1.3	39	6	0.21	7	27	37
EA-BPPO/TEP-1.4%P	1.4	42	20	0.16	12	38	47
EA-BPPO/TEP-1.5%P	1.5	42	3	0.20	15	28	35
tBuA-BPPO/TEP-1.0%P	1.0	39	19	0.14	11	31	41
AM-BPPO/TEP-1.0%P	1.0	51	13	0.18	11	33	40
AM-BPPO/TEP-1.3%P	1.3	50	12	0.12	13	37	45
AM-BPPO/TEP-1.5%P	1.5	57	3	0.17	20	50	50
PA-BPPO/TEP-1.0%P	1.0	39	16	0.18	12	32	40
SU-BPPO/TEP-1.0%P	1.0	50	23	0.23	12	70	42
DPF-BPPO/TEP-1.0%P ^a	1.0	-	-	-	15	>200	51
DMI-BPPO/TEP-1.0%P	1.0	35	16	0.16	11	37	46
DMI-BPPO/TEP-1.5%P	1.5	47	13	0.15	10	30	41
HQ-BPPO/TEP-1.0%P	1.0	55	8	0.19	10	15	15
HQ-BPPO/TEP-1.5%P	1.5	58	22	0.33	5	15	15
HP-BPPO/TEP-1.0%P	1.0	39	14	0.19	6	22	29
DMI-DOPO/TEP-1.0%P	1.0	35	11	0.14	12	31	40
DMI-DOPO/TEP-1.5%P	1.5	42	19	0.20	13	44	50
HQ-DOPO/TEP-1.0%P	1.0	39	2	0.19	8	23	30
HQ-DOPO/TEP-1.5%P	1.5	45	23	0.19	5	16	16

^a Yielded unstable foam.

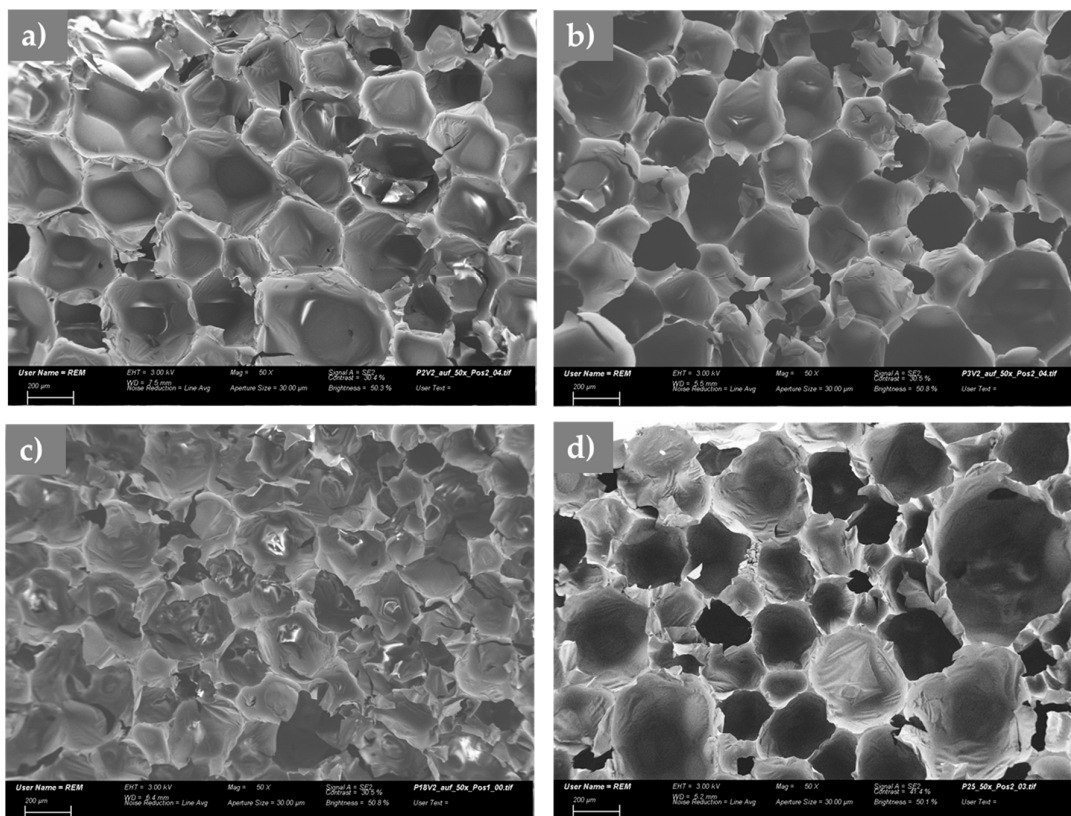


Figure 3. Examples of SEM images of foam (a) TEP-0.3%P; (b) TPP/TEP-1.0%P; (c) MA-BPPO/TEP-1.0%P; and (d) AM-BPPO/TEP-1.3%P, scale bar: 200 µm.

The water uptake WAV is a measure for cell integrity (open or partially perforated cell windows). The values that are summarized in Table 3 indicate different degrees of cell integrity. The foams with EA-BPPO, MA-BPPO, and HQ-BPPO did not show a change in cell integrity when compared to the control and benchmark foams. All other foams had much larger contents of window perforation.

The mechanical behavior of the selected foams was examined by the compression test until a maximum compression of 60% (Table 4, Figure 4 and Figure S4). In the control foam Ref-0%P, the maximum force F_{max} of approximately 270 kPa was reached at low compression. The addition of TEP in TEP-0.3%P resulted in an increase of F_{max} to about 350 kPa. The foam TPP-0.7%P without TEP, but with TPP showed, as expected, a lower maximum force that was comparable to Ref-0%P. Substitution of TPP by the BPPO-containing FR additives induced slight changes to the compression strength as compared to TPP/TEP-1.0%P with EA-BPPO/TEP-1.0%P having F_{max} at slightly higher value and MA-BPPO/TEP-1.0%P having F_{max} at a slightly lower value. It seems to be surprising that a plasticizer, like TEP, led to higher compression strength. This effect can be explained by the reduction of the glass transition temperature T_g by TEP with reaction time ($T_g = f(t_{reaction})$). The improvement in mobility of polymer chains during the polymerization led to increased levels of diffusion-controlled trimerization. Higher degrees of trimerization induce higher crosslinking densities, and this is visible in the compression strength.

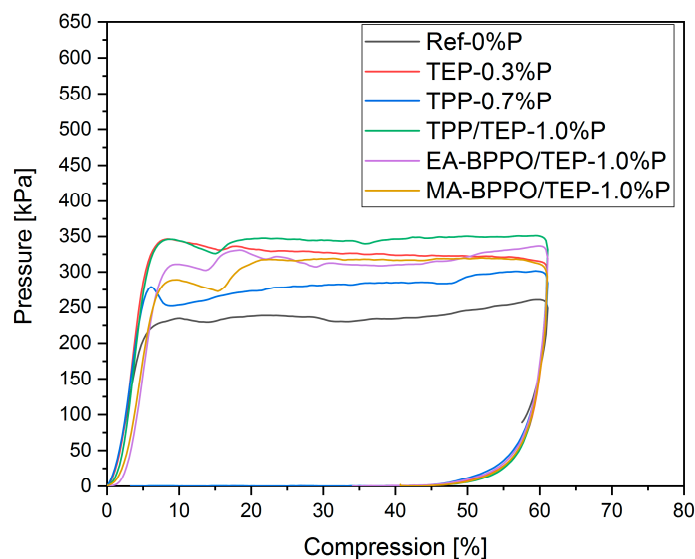


Figure 4. Pressure compression diagram of selected foams.

Table 4. Mechanical properties of selected foams.

Foam	P-content	F_{max}	$F_{60\%}$ ^a
	[wt.%]	[kPa]	[kPa]
Ref-0%P	0	271	197
TEP-0.3%P	0.3	347	255
TPP/TEP-1.0%P	1.0	353	261
TPP-0.7%P	0.7	282	184
EA-BPPO/TEP-1.0%P	1.0	363	234
MA-BPPO/TEP-1.0%P	1.0	314	245

^a Value before unloading at $\epsilon = 60\%$.

3.3. Thermal Decomposition of the Foams

The thermal decomposition of the foams was investigated by TGA under nitrogen up to 800 °C, in combination with pyrolysis-GC/MS executed at the decomposition maxima found in TGA. Table 5 summarizes the results together with a comparison of the residues that are found at the end of the cone calorimeter test (Section 2.4.2.). The pyrolysis products that evolved in the gas phase and were observed by pyrolysis-GC/MS are summarized in Table S4 in the Supplementary Materials. The TGA curves in Figure 5 illustrate the decomposition.

The addition of TEP in TEP-0.3%P and TPP/TEP-1.0%P (Figure 5a) did not alter the main decomposition maximum (3) in the temperature range of 320 °C, but induced a new maximum (1) at a lower temperature (192 °C). This maximum is connected to the vaporization of TEP, as reported previously [2] and proven by pyrolysis-GC/MS (Supplementary Materials, Table S4). Consequently, it was found in all samples containing TEP.

The addition of TPP caused a new maximum (2) at 216 °C that is connected to the evaporation of TPP [39–41], which was proven by pyrolysis-GC/MS. The flame-retardant effect of both TEP and TPP originates from their vaporization at temperatures that are lower than T_{max} (3). The residues of Ref-0%P, TEP-0.3%P, and TPP/TEP-1.0%P found at 800 °C were comparable. These observations suggest that both TPP and TEP completely evaporate in the gas phase.

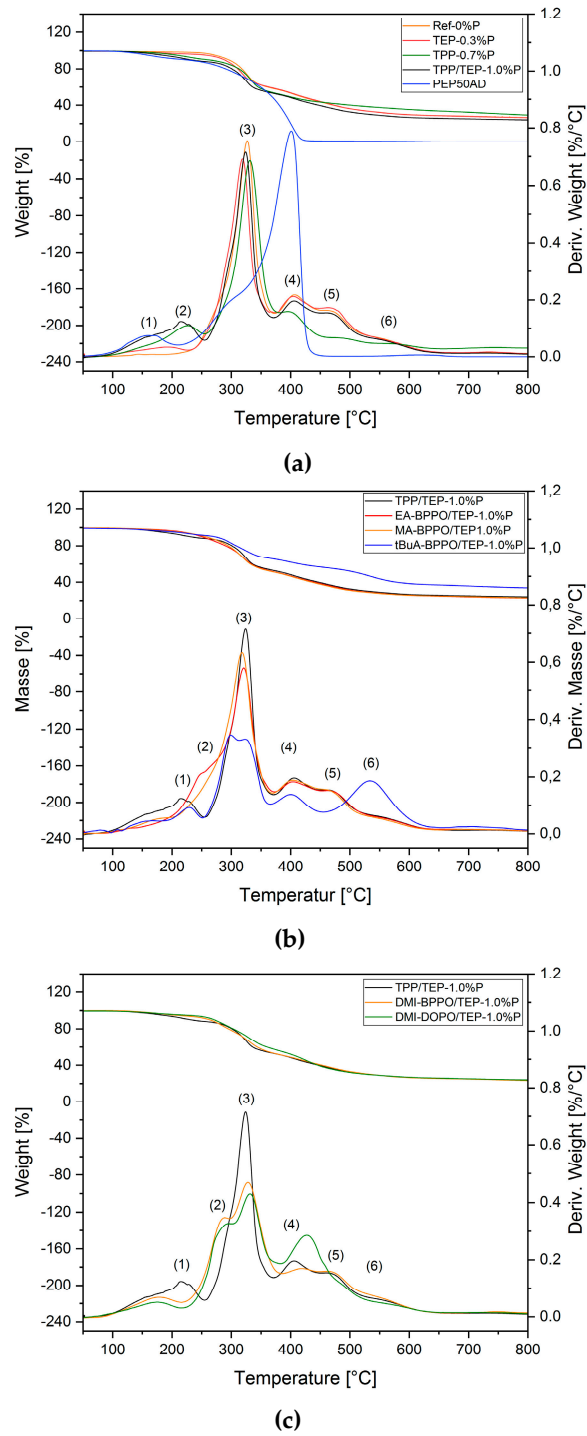


Figure 5. Thermogravimetric analysis (TGA) results: (a) TGA curves of benchmark foams and polyol used; (b) TGA curves of the benchmark foam compared to foams containing BPPO-derivatives; and, (c) TGA curves of the benchmark foam compared with a foam with BPPO-derivative and a DOPO-derivative. All the formulations shown here are specified in Table S1 in the Supporting Information.

Table 5. Decomposition characteristics of the studied foams (description of methods: see Experimental part).

Foam	P-Content	TGA T_{max}	TGA Residue	Fire Residue
	[wt.%]	[°C]	[wt.%]	[wt.%]
Ref-0%P	0.0	326	26.6	22.3
TEP-0.3%P	0.3	319	26.4	26.7
TPP-0.7%P	0.7	330	29.3	44.7
TPP/TEP-1.0%P	1.0	324	24.0	36.4
TPP/TEP-1.3%P	1.3	328	23.8	27.8
TPP/TEP-1.5%P	1.5	331	19.6	28.6
MA-BPPO/TEP-1.0%P	1.0	318	22.9	28.3
MA-BPPO/TEP-1.2%P	1.2	320	24.6	33.1
MA-BPPO/TEP-1.4%P	1.4	315	25.0	34.1
EA-BPPO-0.7%P	0.7	331	24.6	28.2
EA-BPPO/TEP-1.0%P	1.0	321	22.6	28.6
EA-BPPO/TEP-1.3%P	1.3	314	21.3	29.2
EA-BPPO/TEP-1.4%P	1.4	329	22.0	-
EA-BPPO/TEP-1.5%P	1.5	319	23.7	35.7
tBuA-BPPO/TEP-1.0%P	1.0	308	33.4	27.0
AM-BPPO/TEP-1.0%P	1.0	316	30.0	-
AM-BPPO/TEP-1.3%P	1.3	295	31.2	-
AM-BPPO/TEP-1.5%P	1.5	319	30.3	-
PA-BPPO/TEP-1.0%P	1.0	320	34.6	-
SU-BPPO/TEP-1.0%P	1.0	334	28.5	-
DMI-BPPO/TEP-1.0%P	1.0	334	29.1	33.2
DMI-BPPO/TEP-1.5%P	1.5	319	23.6	-
HQ-BPPO/TEP-1.0%P	1.0	334	33.5	31.7
HQ-BPPO/TEP-1.5%P	1.5	315	28.9	-
HP-BPPO/TEP-1.0%P	1.0	321	28.1	-
DMI-DOPO/TEP-1.0%P	1.0	331	23.9	29.4
DMI-DOPO/TEP-1.5%P	1.5	278	24.5	-
HQ-DOPO/TEP-1.0%P	1.0	325	23.3	38.0
HQ-DOPO/TEP-1.5%P	1.5	327	23.8	-

Figure 5b illustrates the thermal decomposition of foams, where the acrylate-BPPO additives substituted TPP. When EA-BPPO was used instead of TPP, a slight change of the TGA curve was visible; the shift of maximum (2) to higher temperature (248 °C) was attributed to the vaporization of EA-BPPO or BPPO fragments. The structurally comparable MA-BPPO caused further changes. The shift of maximum (2) towards maximum (3) and of maximum (3) to a lower temperature was observed, as well as the disappearance of maximum (5). tBuA-BPPO led to the occurrence of a double peak. The structure of the t-butyl group additionally resulted in a strong enhancement of maximum (6) and induced charring. This correlated to a significantly higher residue at 800 °C (char increased by 10 wt.% as compared to TPP/TEP-1.0%P).

The pyr-GC/MS spectra of all the samples taken around 300 °C, i.e., at decomposition maximum (3) indicated the formation of diethylene glycol and phthalic acid anhydride or phthalic acid, respectively. These products are expected for the used polyester polyol co-monomer (see chemical structure of the polyester polyol in Figure 2). All other decomposition compounds were found in minor concentrations. Table S4 summarizes all of the decomposition products that were found in the samples studied for the respective temperature; Figures S5–S10 show the mass spectra that were obtained.

3.4. Fire Behavior of the Foams

3.4.1. Vertical Flame Spread

Table 6 summarizes the flame length observed in the VFS test. The control foam Ref-0%P without phosphorus (TEP or TPP) completely burned. The addition of TEP in TEP-0.3%P led to a slight reduction of the flame length. Adding the benchmark FR TPP, the foam TPP/TEP-1.0%P could have passed the “B2” qualification with a VFS of 15 cm.

Table 6. Results of vertical flame spread vertical flame spread (VFS) test (DIN 4102).

Foam	P-content [wt.%]	Density [kg·m ⁻³]	VFS [cm]
Ref-0%P	0.0	36	>20
TEP-0.3%P	0.3	36	18
TPP-0.7%P	0.7	36	15
TPP/TEP-1.0%P	1.0	39	14
TPP/TEP-1.3%P	1.3	40	13
TPP/TEP-1.5%P	1.5	39	12
MA-BPPO/TEP-1.0%P	1.0	35	15
MA-BPPO/TEP-1.2%P	1.2	47	14
MA-BPPO/TEP-1.4%P	1.4	35	13
EA-BPPO-0.7%P	0.7	36	17
EA-BPPO/TEP-1.0%P	1.0	42	16
EA-BPPO/TEP-1.3%P	1.3	38	13
EA-BPPO/TEP-1.4%P	1.4	39	13
EA-BPPO/TEP-1.5%P	1.5	37	<10
tBuA-BPPO/TEP-1.0%P	1.0	37	15
AM-BPPO/TEP-1.0%P	1.0	51	18
AM-BPPO/TEP-1.3%P	1.3	50	18
AM-BPPO/TEP-1.5%P	1.5	57	18
PA-BPPO/TEP-1.0%P	1.0	39	11
SU-BPPO/TEP-1.0%P	1.0	50	18
DMI-BPPO/TEP-1.0%P	1.0	39	15
DMI-BPPO/TEP-1.5%P	1.5	42	<10
HQ-BPPO/TEP-1.0%P	1.0	55	18
HQ-BPPO/TEP-1.5%P	1.5	58	18
HP-BPPO/TEP-1.0%P	1.0	39	17
DMI-DOPO/TEP-1.0%P	1.0	42	14
DMI-DOPO/TEP-1.5%P	1.5	39	16
HQ-DOPO/TEP-1.0%P	1.0	39	<10
HQ-DOPO/TEP-1.5%P	1.5	45	<10

Foams with nearly all BPPO-acrylates and itaconates passed the VFS-specification of DIN 4102 at a phosphorus concentration of 1 wt.%. Increasing the P-concentration further reduced the flame spread. A comparison of BPPO-itaconate containing foams with DMI-DOPO containing foams showed that the latter had a lower efficiency. Formulations with higher P-contents could not be properly homogenized and they yielded heterogeneous foams. The use of the hydroquinone additive HQ-BPPO led to foams that did not pass the VFS specification, while HQ-DOPO led to a very low flame spread. The foams that were prepared with the amide AM-BPPO yielded flame lengths higher than 15 cm, thus not passing the VFS specification. Even if the P-content was increased, the test could not be passed with AM-BPPO. One of the reasons might be the different density of these foams. A second explanation could be that the hydroquinones and amides are not

inert as solids, but rather react with isocyanate to form urethane and acylurea, leading to lower NCO/OH ratios than the other foams—although NMR could not prove this reaction. It is known that both aromatic urethanes and acylureas are unstable beyond 120 °C, adding further complexity to the polymerization reaction.

The VFS test yields results that give an indication of FR activity. More detailed results are provided by cone calorimetry under forced flaming conditions, as discussed in the next section.

3.4.2. Forced Flaming Combustion

Foams with EA-BPPO, MA-BPPO, tBuA-BPPO, DMI-BPPO, and DMI-DOPO having a roughly similar density and pore sizes were selected for the examination of the developing fire behavior in forced flaming combustion by cone calorimeter. HQ-BPPO and HQ-DOPO were analyzed, despite the different densities to expand the comparison between BPPO and DOPO additives.

Figure 6 compares the HRR over time of foam Ref-0%P to foams with EA-BPPO, TEP, and TPP. The curve for the control foam Ref-0%P is roughly comparable with that results that were reported in the literature [3], although the pentane levels in the cells here and in the literature differ significantly. The foam TEP-0.3%P showed a reduction of HRR in the steady burning phase between 50 and 275 s. The addition of TPP in TPP/TEP-1.0%P did not alter the first phase of burning until 200 s, but it reduced the HRR after this time. When TPP was substituted by EA-BPPO in EA-BPPO/TEP-1.0%P the pentane-related first peak appeared later, as desired, and it was lower than for the benchmark foam TPP/TEP-1.0%P.

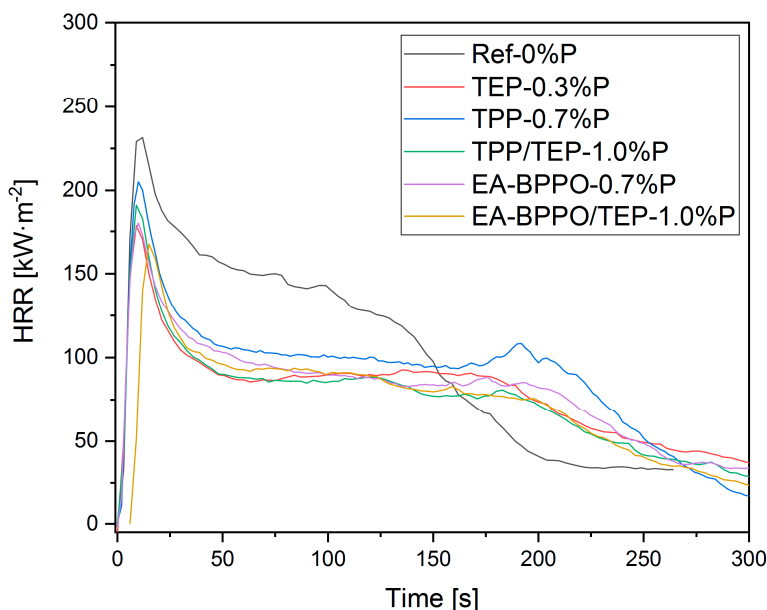


Figure 6. Heat release rate curves (average of three measurements) obtained by forced flaming combustion with heat flux $50 \text{ kW}\cdot\text{m}^{-2}$ of control foams (Ref-0%P, TEP-0.3%P, TPP-0.7%P, and TPP/TEP-1.0%P) as compared to the foams EA-BPPO-0.7%P and EA-BPPO/TEP-1.0%P.

Table 7 summarizes the results of the measurements (selected parameters). The complete data set is given in the SI (Table S6).

Table 7. Results of cone calorimeter tests on the PIR foams under study (heat flux 50 kW·m⁻²).

Foam	PHRR	MARHE	THR	Residue	EHC	TSR
	[kW·m ⁻²]	[kW·m ⁻²]	[MJ·m ⁻²]	[wt.%]	[MJ·kg ⁻¹]	[m ² ·m ⁻²]
Ref-0%P	233	172	24	22.3	2.0	551
TEP-0.3%P	178	128	26	26.7	2.1	392
TPP-0.7%P	205	145	27	44.7	2.1	755
TPP/TEP-1.0%P	164	111	20	36.4	1.8	324
TPP/TEP-1.3%P	163	114	24	27.8	1.8	407
TPP/TEP-1.5%P	162	112	22	28.6	1.8	389
MA-BPPO/TEP-1.0%P	185	132	28	28.3	2.2	527
MA-BPPO/TEP-1.2%P	162	112	23	33.1	2.0	421
MA-BPPO/TEP-1.4%P	167	114	25	34.4	1.9	488
EA-BPPO-0.7%P	180	132	26	28.2	2.0	501
EA-BPPO/TEP-1.0%P	168	121	23	28.6	1.9	433
EA-BPPO/TEP-1.3%P	156	113	24	29.2	1.9	506
EA-BPPO/TEP-1.5%P	139	103	27	35.7	2.1	398
tBuA-BPPO/TEP-1.0%P	166	114	26	27.0	2.0	380
DMI-BPPO/TEP-1.0%P	171	118	21	33.2	1.9	317
DMI-DOPO/TEP-1.0%P	162	108	24	29.4	1.9	371
HQ-BPPO/TEP-1.0%P ^a	374	266	29	31.7	1.7	1228
HQ-DOPO/TEP-1.0%P	134	101	22	38.0	1.8	503

^a Deviating density (50 kg·m⁻³).

All of the foams ignited after 1 s. This is typical for insulating foams with low density that do not dissipate the heat from the heater into the material [42]. The maximum of the heat release rate (PHRR) was reached 10 s after ignition in all of the samples. Fast flame spread after ignition is followed by fast formation of a charring layer, which lowers the heat release. For most foams, only little shrinkage was observed which can be attributed to the crosslinking density of the PIR foams. Figure 7 illustrates the appearance of selected foams after the cone calorimeter test.

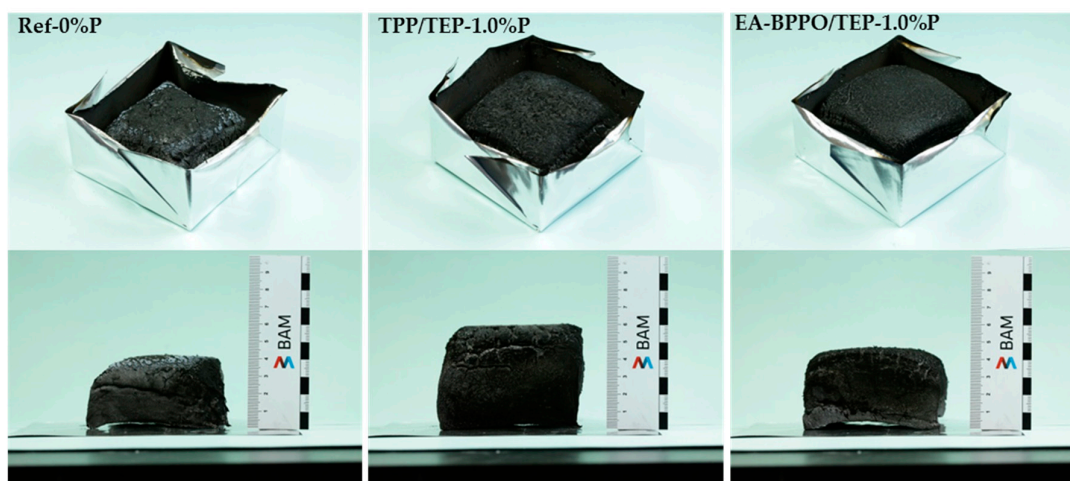


Figure 7. Images of foam Ref-0%P, TPP/TEP-1.0%P, and EA-BPPO/TEP-1.0%P after the cone calorimeter test.

The control foam Ref-0%P left substantial residue with a height of 4 cm after the cone experiment. This illustrates the efficiency of the isocyanurate rings in char formation. The addition of TEP alone in TEP-0.3%P raised the char height to 4.5 cm, and the further addition of condensed phase-active TPP to 6 cm. Exchange of TPP by EA-BPPO in EA-BPPO/TEP-1.0%P altered the char

height to 4.5 cm, a height that was observed in all foams with BPPO additives. Therefore, the chars of foams with BPPO additives had higher density. The phosphorus additives enhanced the residue to about 30 wt.% in each case. Furthermore, increasing the P-content raised the residue accordingly.

The addition of TEP resulted in a drastic decrease in PHRR (TEP-0.3%P). Further addition of TPP in TPP/TEP-1.0%P formulation did not alter the PHRR. The same applies for MARHE. The substitution of TPP by the BPPO derivatives to meet the comparable P-content (1 wt.%) yielded PHRR and MARHE values that were comparable to TPP/TEP-1.0%P. The PHRR values that were found correlated to the P-content, except for MA-BPPO/TEP-1.4%P, Figure 8. THR was hardly influenced by all the additives that we ascribe to the comparable P-contents and the comparable amount of burnt material. The effective heat of combustion was also hardly affected by the type of FR additives.

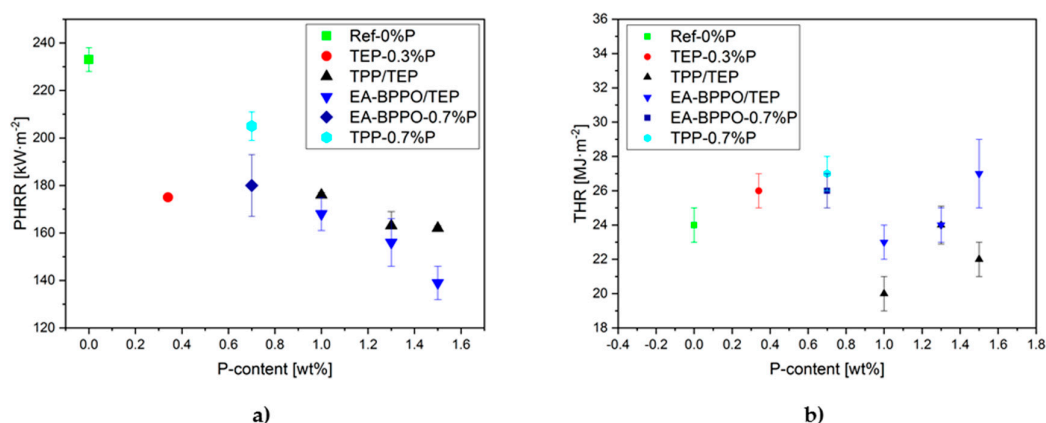


Figure 8. Correlation of (a) peak heat release rate (PHRR) and (b) total heat released (THR) of selected foams to their P-content.

The dimethyl itaconates were used to compare DOPO and BPPO at a phosphorus loading of 1 wt.%. Significant differences between DMI-DOPO (DMI-DOPO/TEP-1.0%P) and DMI-BPPO (DMI-BPPO/TEP-1.0%P) were not found in the cone calorimeter tests. Both performed in the range of TPP/TEP-1.0%P benchmark, except for a lower TSR (Table 7). The TSR of both samples was lower than that with the BPPO-acrylates. The CO yield in the samples with BPPO-based FR, except for the itaconate, was reduced by 10% when compared to the benchmark TPP/TEP-1.0%P (Supplementary Materials, Table S6). The sample HQ-DOPO/TEP-1.0%P showed a clear but reproducible irregularity in the parameters (higher PHRR, MARHE, TSR), which we attribute to the higher density of the foam. Higher density resulted from the altered reaction kinetics (see Section 3.2.).

Figure 9 compares vertical flame spread and the maximum heat release rate. They both follow the same trend: samples with lower PHRR (i.e., with higher P-content) also show lower VFS.

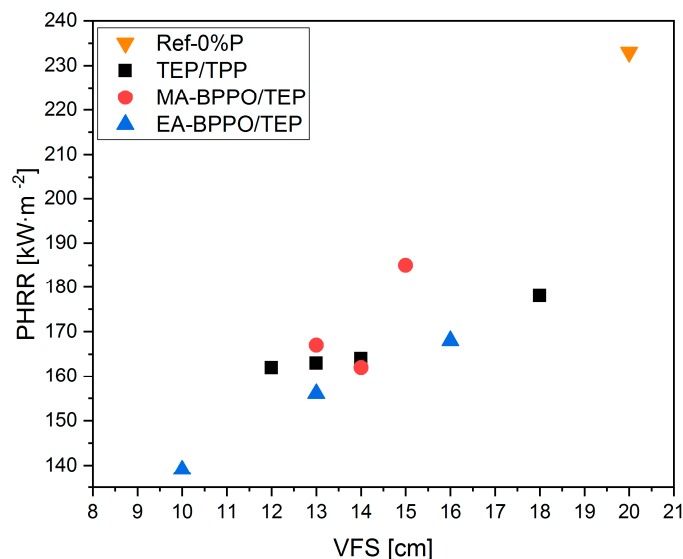


Figure 9. Correlation between PHRR and VFS of the reference foams, MA-BPPO/TEP, and EA-BPPO/TEP.

4. Conclusions

A series of new flame retardants (FR) based on dibenzo[d,f][1,3,2]dioxaphosphepine 6-oxide (BPPO) incorporating acrylates and benzoquinone were tested in PUR/PIR foams with respect to their fire performance. All FR agents were tested at comparable P-concentrations. The majority of the novel BPPO derivatives did not alter the physical properties (density and morphology) of the PIR foams. We found differences in the FR properties (PHRR, MARHE and TSR) based on the substituents of the acrylate employed. At 1 wt% P, the structurally more complex tBuA-BPPO turned out to be the most effective FR additive due to lower PHRR, MARHE, TSR as well as a higher residue compared to the benchmark triphenyl phosphate (phosphorus oxidation state +V). The majority of the BPPO-acrylate and itaconate derivatives behaved similar to the benchmark, triphenyl phosphate. The novel FR were also compared to chemically similar substances based on 9,10-dihydro-9-oxy-10-phosphaphenanthrene-10-oxide (DOPO) with phosphorus in the oxidation state -I to investigate the influence of oxidation state on gas phase and condensed phase contributions to flame retardancy in PUR/PIR foams. In forced flaming combustion experiments a significant influence of the FR on charring, TSR, and MARHE was found. An enhanced charring behavior for the P(+III) compounds compared to P(-I) could be confirmed according to statements in literature [2]. This agrees with the higher gas phase activity of P(+I) derivatives observed. In summary, the BPPO-containing foams combine fire protection with increased compression strength. These findings have practical relevance for tuning FR properties of PIR foams to meet safety and regulatory requirements, e.g., in the building and construction industry.

Supplementary Materials: The following supplementary materials are available online at www.mdpi.com/xxx/s1, Figure S1. ATR-FTIR spectra of selected foams, Figure S2. FTIR spectra of pure EA-BPPO, Figure S3. ATR-FTIR spectra of foam Ref-0%P and EA-BPPO/TEP-1.0%P compared with pure EA-BPPO, Figure S4. Stress-strain diagrams of the foams selected foams, Figure S5. Pyr-GC/MS spectra of Ref-0%P, Figure S6. Pyr-GC/MS spectra of TEP-0.3%P, Figure S7. Pyr-GC/MS spectra of TPP/TEP-1.0%P, Figure S8. Pyr-GC/MS spectra of EA-BPPO/TEP-1.0%P, Figure S9. Pyr-GC/MS spectra of DMI-BPPO/TEP-1.0%P, Figure S10. Pyr-GC/MS spectra of DMI-DOPO/TEP-1.0%P, Table S1. Foam compositions studied, Table S2. Polycondensation experiments of SU-BPPO with 1,4-butanediol, Table S3. Transesterification experiments with EA-BPPO and 1,4-butanediol, Table S4. Decomposition products observed with Pyr-GC/MS, Table S5. Quantitative P-contents of selected foams and their residues after cone calorimetry, Table S6. Complete results of cone calorimeter tests on the PIR foams under study.

Author Contributions: J.L. carried out specimen preparation, experiments and data evaluation. J.L. performed the experiments and wrote this article together with D.P. D.P. acquired the project and developed the project idea, supervised the work and wrote this article. M.P. and R.W.A. contributed to the scientific discussion and collaborated on the writing of this publication. R.W.A. developed the project idea. M.G. performed cone calorimeter experiments, evaluated the data and contributed to the scientific discussion of the results. B.S. contributed to the scientific discussion of the results and in writing the paper. B.V. supervised the PhD work of J.L.

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